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Stoichiometry-KEY Balance the following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$ 2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ 3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$ 4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$ 5) $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3$ Using the following equation: $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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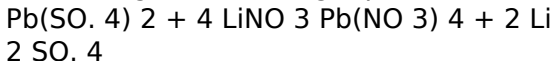
Balancing Equations and Simple Stoichiometry-KEY

Mole Conversions and Stoichiometry
Review Worksheet. 1) Using the following
equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
How many grams of sodium
sulfate will be formed if you start with
200 grams of sodium hydroxide and you

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Answers

have an excess of sulfuric acid (H. 2 SO
4)?2)Using the following equation:



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Worksheet. 1) Using the following
equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} +$

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Answers

Na₂SO₄. 4 How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an

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as some people pronounce it, "stoi-shee-oh-met-tree." Don't pronounce it that way, it'll make you sound silly. The actual pronunciation: "stoy-key-ah-meh-tree." Now that we've got that out of the way, let's learn about the magical world of stoichiometry! The magical world of stoichiometry In...

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