

# Specific Heat Worksheet 2 Answers

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### Specific Heat Worksheet 2 Answers

$q$  = amount of heat (J)  $m$  = mass (grams)  $c$  = specific heat (J/g°C)  $\Delta T$  = change in temperature (°C)  
2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. Heat is a combination of kinetic energy (measured by temperature) and potential energy. a. Perform calculations using: ( $q = m c \Delta T$ ) b.

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## Worksheet- Calculations involving Specific Heat

Specific Heat Worksheet #2 (Answer Key) 1. Brass is an alloy made from copper and zinc. A 0.66 kg sample of brass at 98.6 o C is dropped into 2.33 kg of water at 4.6 o C.

## Specific Heat Worksheet-2Answer Key - Specific Heat ...

Name Answer Key Date 9/9/15 Chp 2-1: Specific Heat Worksheet (m) ( $\Delta T$ ) (C sp )=Q 1. Specific heat is the amount of energy that it takes to raise the temperature of 1 gram of a substance by 1 degree kelvin 2. Absolute zero is the temperature at which all molecular motion ceases 3. Endothermic process is a change in matter in which energy is absorbed 4.

## Specific Heat WS Answers - Name Answer Key Date Chp 2-1 ...

Specific Heat Problems Worksheet Answers Worksheet December 25, 2018 03:29 To be able to properly identify what kind of heating and cooling problem you are having, you will need to refer to a Worksheet Answers to Heat and Cooling Problems.

## Specific Heat Problems Worksheet Answers

Before discussing Calculating Specific Heat Worksheet Answers, you need to recognize that Knowledge can be your answer to a better the next day, along with studying doesn't just stop the moment the school bell rings. Of which getting claimed, many of us provide you with a a number of basic yet helpful posts along with design templates made ideal for almost any educative purpose.

## Calculating Specific Heat Worksheet Answers | akademiexcel.com

Worksheet- Calculations involving Specific Heat 1. For  $q = m \cdot c \cdot \Delta T$ : identify each variables by name & the units associated with it. 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. (-m.c.AT) a. Perform calculations usin 1. Gold has a specific heat of

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0.129 J/(g·°C). How

### North St. Paul-Maplewood Oakdale / Overview

Specific Heat Practice Worksheet 1. An aluminum skillet weighing 1.58 kg is heated on a stove to 173 °C. Suppose the skillet is cooled to room temperature, 23.9 °C. How much heat energy (joules) must be removed to cause this cooling? The specific heat of aluminum is 0.901 J/(g · °C). 2.

### Specific Heat Practice Worksheet

Specific Heat Worksheet Name (in ink):  $C = q/m\Delta T$ , where  $q$  = heat energy,  $m$  = mass, and  $T$  = temperature Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ . Show all work and proper units. Answers are provided at the end of the worksheet without units. 1. A 15.75-g piece of iron sorbs 1086.75 joules of heat energy, and its temperature changes from 25 °C to 175°C.

### Specific Heat Wksht20130116145212867

Chapter 10 Worksheet #2 1. Calculate the energy require (in calories) to heat 10.4 g of mercury from 37.0 °C to 42.0 °C. Specific heat of mercury is 0.14 J/g °C.  $q = m c \Delta t$   $q = 10.4 \text{ g} \cdot 0.14 \text{ J/g } \cdot 5.00 \text{ }^\circ\text{C} = 7.28 \text{ J} \cdot 1 \text{ cal} = 1.74 \text{ cal}$  2. If 50. J of heat are applied to 10. g of iron, by how much will the temperature of the iron

### Chapter 10 Worksheet #2 Answer

Specific Heat DIRECTIONS: Use  $q = (m)(\Delta T)(C_p)$  to solve the following problems. Show all work and units. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C.

### Specific Heat Worksheet

Specific Heat Worksheet Answers from specific heat worksheet answer key ,

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### Specific Heat Worksheet Answer Key - Briefencounters

2d.  $q = (25.0 \text{ g}) (25.0 \text{ }^\circ\text{C}) (2.06 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1})$  3.  $q = (3.21 \text{ g}) (4.0 \text{ }^\circ\text{C}) (4.184 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1})$  4.  $q = (55.6 \text{ g}) (34.9 \text{ }^\circ\text{C}) (4.184 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1})$ . How did I know this was liquid water? 5. Same type as #4, but no answer. Good luck!! 6b.  $(6.02 \text{ kJ/mol}) (74.5 \text{ g} / 18.0 \text{ g/mol})$  6d.  $(40.7 \text{ kJ/mol}) (43.89 \text{ g} / 18.0 \text{ g/mol})$  7. Some hints: a) water means liquid water

### Thermochem Worksheet #2 Answers - ChemTeam

Here are the heat capacities of the four substances:  $0.10 \text{ cal/g }^\circ\text{C}$ ,  $0.25 \text{ cal/g }^\circ\text{C}$ ,  $1.0 \text{ cal/g }^\circ\text{C}$ , &  $0.2 \text{ cal/g }^\circ\text{C}$ . Match & then label each substance with its specific heat capacity on the graph. See graph above. 7. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its temperature? Explain ...

### Worksheet- Introduction to Specific Heat Capacities

from  $25^\circ\text{C}$  to  $175^\circ\text{C}$ . Calculate the specific heat capacity of iron.  $q = m \cdot c \cdot \Delta T$   
 $Q = -2 \cdot f \cdot c \cdot Tr = 12 \cdot f \cdot c \cdot It \cdot C = 7 \cdot Q \cdot \wedge$ ,  $A \cdot cx \cdot 4 / o \cdot J / 4 \cdot c \cdot T - q \cdot L = C \cdot rh [f \cdot l \cdot r \cdot T \cdot 6, 7 \cdot r \cdot l - o; re (nr' - Lr \cdot z)]$ .  $= -c \cdot fl \cdot c \cdot o? "r \cdot 7 \cdot 11$ .  
How many joules of heat are needed to raise the temperature of  $10.0 \text{ g}$  of aluminum from  $22^\circ\text{C}$  to  $55^\circ\text{C}$ , if the specific heat of aluminum  $Q \dots$

### Specific Heat Capacity - Worksheet (key) [d4p7my5q264p]

Heat Transfer/ Specific Heat Problems Worksheet Solving For Heat (q) 1. How many joules of heat are required to raise the temperature of  $550 \text{ g}$  of water from  $12.0 \text{ }^\circ\text{C}$  to  $18.0 \text{ }^\circ\text{C}$ ? 2. How much heat is lost when a  $64 \text{ g}$  piece of copper cools from  $375 \text{ }^\circ\text{C}$ , to  $26 \text{ }^\circ\text{C}$ ? (The specific heat of copper is

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0.38452 J/g x °C). Place your answer in kJ. 3.

### Heat Transfer/ Specific Heat Problems Worksheet

Download Free Chemistry Specific Heat Worksheet Answers Specific Heat Worksheet Name (in ink):  
 $C = q/m\Delta T$ , where  $q$  = heat energy,  $m$  = mass, and  $T$  = temperature Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ . Show all work and proper units. Answers are provided at the end of the worksheet without units. 1. A 15.75-g

### Chemistry Specific Heat Worksheet Answers

$5275 \text{ J} = 50 \text{ g} \times 0.5 \text{ J/g} \times \Delta T$ .  $\Delta T = 5275/50 \times 0.5$ .  $\Delta T = 211$ . So it will rise 211 degrees making final temp  $211 + 20 = 231$  degrees. 100.0mL of 4.0 °C water is heated until its temperature is 37 °C....

### Specific Heat Worksheet? | Yahoo Answers

Specific Heat Calculations Worksheet Name Chemistry 2 points from Specific Heat Worksheet Answers, source:yumpu.com Heat Fusion Worksheet Free Worksheets Library from Specific Heat Worksheet Answers, source:comprar-en-internet.net

### Specific Heat Worksheet Answers | Homeschooldressage.com

Heat Transfer Specific Heat Problems Worksheet - Temperature is a typical value of energy for every one of the molecules and atoms in a particular system. It's an ordinary worth of energy for molecules and all of the atoms in a system that is given. The quantity of energy obtained is equivalent to the total when everything is at the same temperature.

### Heat Transfer Specific Heat Problems Worksheet

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