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This was one of the better texts that I used for a graduate course in chemical hydrogeology.

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AQUEOUS SOLUTION GEOCHEMISTRY

Acid = substance containing hydrogen which gives free hydrogen (H^+) when

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dissolved in water Base = substance containing the OH group that yields free (OH-) when dissolved in water An acid solution is one containing an excess of free H+, and a base is one containing excess of free OH-.

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AQUEOUS AND ENVIRONMENTAL GEOCHEMISTRY

most useful in the context of gases and liquids. A single liquid phase, such as an

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aqueous solution, may contain a number of species. For example, H_2O , H_2CO_3 , HCO_3^- , CO_3^{2-} , H^+ , and OH^- are all species commonly present in natural waters. The term species is generally reserved for an entity

W. M. White Geochemistry Chapter 3: Solutions

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Eby, G. N. (2004) Principles of Environmental Geochemistry. Brooks/Cole, 514 p. A review copy can be obtained from Brooks/Cole.. The solution manual, copies of the Excel spreadsheets used to solve the problems, and PowerPoint files containing all the text tables and figures are available from the publisher for

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Measure the concentrations (molar) of cations and anions in your solution
2nd Step: From the concentrations (c) and ion charges (Z) calculate the Ionic Strength (I)
3th Step: Using the ionic

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[Donald Langmuir] -- This book offers coverage of controls on the chemical quality of surface and subsurface waters, both pristine and polluted, with an emphasis on problem-solving and practical applications.

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