

### Chemquest 17 Answer Key

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$8.4 - 5 - 17 - 22 = x \times 10 = 2.05 \times 10$ . 17.  $4.1 \times 10$ . 4.1. b)  $5.4 \times 10$ .  $7.3 \times 10$ . 32.  $14 = 5.4$ .  $7.3 \times 10$ .  $32 - 14 = .74 \times 10$ . 18 =  $7.4 \times 10$ . 17. Information: Adding and Subtracting Using Scientific Notation.

Whenever you add or subtract two numbers in scientific notation, you must make sure that they have the same exponents. Your answer will then have the ...

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17. What relationship exists between a homogeneous mixture and the number of phases in the mixture? A homogeneous mixture must have only one phase. 18. What is the difference between

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homogeneous and heterogeneous mixtures? A homogeneous mixture is the same throughout—one small sample of the mixture is an exact representation of the whole.

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17. What is wrong with the following notation? You should find two things wrong. There is no such thing as a d sublevel in the 2nd energy level. Also, a d sublevel needs to have 5 lines drawn for the 5 orbitals in the d sublevel. 18. Write the notation for an electron in the 4th energy level in an f sublevel spinning clockwise.

### **Information: Energy Levels and Sublevels**

b) 32 14 18 17 14 32 10 .74 10 7.4 10 7.3 5.4 7.3 10 5.4 10 x x x x x = - = = Information: Adding and Subtracting Using Scientific Notation Whenever you add or subtract two numbers in scientific notation, you must make sure that they have the same exponents. Your answer will then have the same exponent as the numbers you add or subtract. Here ...

### **Information: Qualitative vs. Quantitative**

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3. Iodine has 17 more electrons than krypton and so you can begin by writing [Kr] followed by orbitals for 617 more electrons. After 4p comes 5s2... Iodine = [Kr]5s24d105p5 Any of the noble gases can be used for abbreviations. Here are a few more examples: iron = [Ar]4s23d6 cesium = [Xe]6s1 phosphorus = [Ne]3s23p3

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Page 1 and 2: 1 ChemQuest 1 Name: \_\_\_\_\_ Page 3 and 4: 3 Another example:  $7.45 \times 10^{-8} = 0$  Page 5 and 6: 5 ChemQuest 2 Name: \_\_\_\_\_ Page 7 and 8: 7 6.Round the following numbers so; Page 9 and 10: 9 Example #1: Adding The tens place Page 11 and 12: 11 Information: Elements, Compounds Page 13 and 14: 13 Information: Homogeneous and Het ...

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7; since chlorine is in column 17 (or "7A") it has 7 valence electrons. 3. Since  $\text{CCl}_4$  is made up of one carbon and four chlorine atoms, how many total valence electrons does  $\text{CCl}_4$  have? Hint: add your answer to question 1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$ . Verify that there are 32 electrons pictured in the ...

#### **ChemQuest 23 : Drawing Covalent Compounds**

Calculate the average atomic masses. Round all answers to two decimal places. Record your answer with the correct units. 1. Iodine is 80% 127.1 and 17% 126.1, and 3% 128.1. Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that have the following

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