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moles(grams mole) = grams. Conversely, to convert the mass of a substance to moles, we use. moles(grams mole) = grams . (mass molar mass) → moles . (grams grams / mole) = grams(mole grams) = moles . Be sure to pay attention to the units when converting between mass and moles.

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312 Chapter 11 The Mole Converting Number of Representative Particles to Moles Zinc is used as a corrosion-resistant coating on iron and steel. It is also an essential trace element in your diet. Calculate the number of moles that contain 4.50 10²⁴ atoms of zinc (Zn). 1. You are given the number of atoms of zinc and must find the equiva-lent number of moles.

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answer will provide you with the number needed to . multiply the empirical formula by to get the . molecular formula. Title: Chapter 11 - The Mole Author: WCS Last modified by: WCS Created Date: 3/2/2011 7:16:00 PM Company: WCS Other titles: Chapter 11 - The Mole ...

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Mole 1.. Idenü\$- and calculaterthe number of.representa- five particles. in each of the following quantities. a. 2.15 moles of gold l. b. 0.151 mole ofniü.ogen oxide q.oqxjo NO c. 11.5 moles of potassium bromideG 2. Calculate the number of moles of thè substance that contains the following number of represen- tative parficles.

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This means that 100.00 g of sucrose always contains 42.11 g of carbon, 6.48 g of hydrogen, and 51.41 g of oxygen. First we will use the molecular formula of sucrose (C 12 H 22 O 11) to calculate the mass percentage of the component elements; then we will show how mass percentages can be used to determine an empirical formula.

Chapter 11.2: Empirical and Molecular Formulas - Chemistry ...

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CHAPTER 11 SOLUTIONS MANUAL Section 11.1 Defining Stoichiometry pages 368-372 Practice Problems pages 371-372 1. Interpret the following balanced chemical equa-tions in terms of particles, moles, and mass. Show that the law of conservation of mass is observed. a. N 2 (g) 3H 2 (g) → 2NH 3 (g) 1 molecule N 2 3 molecules H 2 → 2 molecules NH 3 1 mole N 2 3 moles H

StoichiometryStoichiometry

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