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Model 1 — Acid

Strength and

Conductivity HCl

Strong acid Water

molecule Acid molecule

$\text{HCl(g)} + \text{H}_2\text{O(l)}$

$\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ I-IF

Weak acid $\text{HF(g)} +$

$\text{H}_2\text{O(l)}$ $\text{H}_3\text{O}^+(\text{aq}) +$

$\text{F}^-(\text{aq})$ Note: Excess ...

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There are only a few (6) strong acids, so many people choose to memorize them. All the other acids are weak. The strong acids are hydrochloric acid, nitric acid, sulfuric acid, hydrobromic acid, hydroiodic acid, perchloric acid, and chloric acid.

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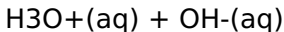
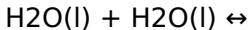
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strong acids include

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HCl, HBr, HI, HClO₄, H₂SO₄, and HNO₃ - molecules of strong acids completely ionize in solution to produce hydronium ions; 100% of molecules of strong acid are ionized in a solution (assuming concentration is not extremely high) - As such, concentration of H₃O⁺ in strong acid solutions are equal to initial concentration of strong acid pH of strong acid solution is

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