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Covalent Bond 1. The hybridization of  
the central carbon in  $\text{CH}_3\text{C}\equiv\text{N}$  and the  
bond angle  $\text{CCN}$  are a.  $\text{sp}^2$ ,  $180^\circ$ . b.  $\text{sp}$ ,  
 $180^\circ$ . c.  $\text{sp}^2$ ,  $120^\circ$ . d.  $\text{sp}^3$ ,  $109^\circ$ . 2.

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Which of the following statements about an  $sp$  hybridized carbon is FALSE? a. It is divalent. b. It forms bonds that are linear. c. It has two  $p$  orbitals. d.

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